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limits. A solution consisting of two components is called binary solution. The component which is present in large quantity is called solvent and the component which is small in quantity is called solute. If both components are in same physical state.

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askITians

A solution is a homogeneous mixture of two or more components in which the particle size is smaller than 1 nm. Common examples of solutions are the sugar in water and salt in water

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solutions, soda water, etc. In a solution, all the components appear as a single phase. There is particle homogeneity i.e. particles are evenly distributed.

Solution - Definition, Properties, Types, Videos & Examples

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Chemistry 108 Lecture Notes Chapter 7: Solutions 3 Solutions • The primary ingredient in a solution is called the \_\_\_\_\_. • The other ingredients are the \_\_\_\_\_ and are said to be dissolved in the solvent. • Water is the most common solvent. • Water is a unique solvent because so



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many substances can dissolve in it.

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2 Solutions Solution is a homogeneous mixture of two or more substances in same or different physical phases. The substances forming the solution are called components of the

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Solution is a homogeneous mixture of two or more substances in same or different physical phases. The substances forming the solution are called components of the solution. On the basis of number of components a solution of two components is called

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binary solution. Solute and Solvent In a binary solution, solvent is the component which is present in large quantity while the other

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1. A solution is a homogeneous

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mixture of two or 9. more chemically non-reacting substances. The components of a solution generally cannot be separated by filtration, settling or centrifuging. 2. A solution may be classified as solid, liquid or a gaseous solution. 3.

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The minor component of the solution is called the solute. In the present example, sugar is the solute. The major component of the solution is called the solvent. In this case water is the solvent. Solutions can also be formed by mixing together many



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different phases of matter. For instance, air is a solution.

Composition of Solutions: Solution Composition | SparkNotes

A chemical solution exhibits several properties: A solution consists of a homogeneous mixture. A solution is

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composed of one phase (e.g., solid, liquid, gas). Particles in a solution are not visible to the naked eye.

Solution Definition in Chemistry -  
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solvent. It is defined as a homogeneous mixture of two or more than two substances. They can be classified into three types: Solid solutions. Gaseous solutions. Liquid solutions. Molarity, mole fraction, percentages, and molality are the terms used to express the

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concentration of a solution.

CBSE Class 12 Chemistry Notes

Chapter 2 Solutions

Chemistry. From aluminum to xenon, we explain the properties and composition of the substances that make up all matter.

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What makes up a solution? 

- A solution has two components: the solute and the solvent.

- The solvent is the substance in greater amount (it does

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the dissolving).

The solute is considered to be the dissolved substance. Chemistry 30S  
Unit 4 - Solutions

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Solutions which contain two  
components in it are called Binary

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Solutions. Substances which are used to prepare a solution are called as Components. The component that is present in the largest quantity is known as Solvent. Solvent determines the physical state in which solution exists.

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Solutions. Solution is a homogeneous mixture of two or more substances in

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same or different physical phases. The substances forming the solution are called components of the solution. On the basis of number of components a solution of two components is called binary solution.

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